

## DESCRIPTION

## CATALYST FOR DIMETHYL CARBONATE SYNTHESIS

## 5 Technical Field

The present invention relates to a catalyst for dimethyl carbonate synthesis used to synthesize dimethyl carbonate. More particularly, it relates to a catalyst for dimethyl carbonate synthesis which can be handled  
10 easily and which is expected to have a high conversion rate.

## Background Art

Dimethyl carbonate (hereinafter, also referred to as  
15 DMC) is a very useful compound as a fuel additive for increasing the octane value of gasoline etc., as a raw material for polycarbonate, which is an engineering plastic, as a carbonylating agent, or the like. Further, the use of dimethyl carbonate as a raw material for a fuel  
20 cell is anticipated. Therefore, the synthesis of dimethyl carbonate has been attempted by using various methods.

As a method for synthesizing dimethyl carbonate, there has conventionally been known a method for obtaining dimethyl carbonate by means of a reaction under a high  
25 pressure condition using carbon dioxide and methanol as

raw materials. However, because of water yielded at the time of reaction, this method has a problem of low yields and selectivity.

Also, as an industrial manufacturing method, there  
5 has generally been known a process for obtaining dimethyl carbonate by means of an esterification reaction using phosgene, which is a highly toxic organic chlorine compound, or carbon monoxide (CO) and methanol as raw materials.

10 On the other hand, in recent years, a synthesis scheme has been developed for obtaining dimethyl carbonate by using CO<sub>2</sub> in a supercritical state and acetone dimethyl acetal (also referred to as 2,2-dimethoxypropane) as raw materials. In such a synthesis scheme, CO<sub>2</sub> in a  
15 supercritical state is allowed to react with acetone dimethyl acetal over a dibutyltin catalyst under conditions of 30 MPa and 180°C. However, since the dibutyltin catalyst is highly toxic and relatively high in cost, and also because the reaction conditions are tough,  
20 it cannot be said that this method is industrially favorable. Furthermore, the yield of dimethyl carbonate in this reaction is about 5%, so that this method is disadvantageous in terms of economy.

As another conventional art, Japanese Patent  
25 Provisional Publication No. 6-025104 (No. 025104/1994) has

disclosed a method in which dimethyl carbonate is produced continuously by a gas-phase reaction in the presence of a solid catalyst consisting of a platinum-group metal or a compound thereof by using carbon monoxide and methyl  
5 nitrite as raw materials. However, this method has a problem of inferior handleability because it uses toxic CO as a raw material.

Japanese Patent Provisional Publication No. 7-069995 (No. 069995/1995) has disclosed a method in which dimethyl  
10 carbonate is produced by continuously subjecting carbon monoxide and methyl nitrite to a gas-phase reaction over a heterogeneous catalyst containing a platinum-group metal. However, because this method also uses toxic CO as a raw material, it has a problem of inferior handleability.

15 Also, Japanese Patent Provisional Publication No. 10-036297 (No. 036297/1998) has disclosed a method in which dimethyl carbonate is produced by transesterification using ethylene carbonate and ethanol. However, since the reaction selectivity is greatly  
20 affected by process conditions such as reaction temperature, this method has a problem in that a large-scale system for strictly controlling the reaction conditions is needed.

25 Disclosure of the Invention

The present invention has been made to solve the above problems, and accordingly an object thereof is to provide an economical catalyst for dimethyl carbonate synthesis which has a high conversion rate and can be  
5 handled easily, the catalyst being used to obtain dimethyl carbonate from acetone dimethyl acetal and  $\text{CO}_2$  in a supercritical state.

The present invention provides a catalyst for dimethyl carbonate synthesis, which is used for producing  
10 dimethyl carbonate from acetone dimethyl acetal and  $\text{CO}_2$  in a supercritical state, wherein the catalyst is obtained by loading a strong acid on a carrier composed of a compound having a solid state.

The catalyst for dimethyl carbonate synthesis is  
15 used favorably in the industrial field because it is low in cost and easy to handle, and has a far higher yield of dimethyl carbonate than that of the conventional method. Also, a system for obtaining dimethyl carbonate from methanol and  $\text{CO}_2$ , which is carried out by using this  
20 catalyst for dimethyl carbonate synthesis, can be applied to the mass production of dimethyl carbonate because it is clean without discharging unnecessary by-products etc., and has high efficiency.

25 Best Mode for Carrying Out the Invention

An embodiment of the present invention will now be described in detail. However, the present invention is not limited to the embodiment described below.

Hereunder, the present invention will be explained  
5 in detail with reference to an embodiment.

According to one embodiment of the present invention, there is provided a catalyst for dimethyl carbonate synthesis which is obtained by loading a strong acid on a carrier composed of a compound having a solid acid site.  
10 In a reaction for obtaining dimethyl carbonate from acetone dimethyl acetal and  $\text{CO}_2$ , a catalyst is required that is capable of dehydrating methanol by activated adsorption of  $\text{CO}_2$ .

The phrase "compound having a solid acid site" means  
15 various solids in which a solid acid carries out acid-base catalysis, such as metal oxides ( $\text{Al}_2\text{O}_3$ ,  $\text{V}_2\text{O}_5$ , etc.), sulfides ( $\text{ZnS}$ , etc.), sulfates ( $\text{NiSO}_4$ ,  $\text{CuSO}_4$ , etc.), phosphates ( $\text{AlPO}_4$ , Ti phosphate, etc.), chlorides ( $\text{AlCl}_3$ ,  $\text{CuCl}_2$ , etc.), clay minerals, and zeolite. In this  
20 embodiment, the use of  $\text{ZrO}_2$ ,  $\text{Al}_2\text{O}_3$ , and  $\text{TiO}_2$  is especially preferable. This is because each of these compounds has a great specific surface area and a proper acid site, and therefore is effective in stably loading  $\text{SO}_4^{2-}$  or  $\text{PO}_4^{3-}$ .  
When these compounds are used as a carrier, they can be  
25 used singly or by mixing two or more compounds.

Further, among the compounds having a solid acid site, a compound having a great specific surface area is preferably used. This is because the compound having a great specific surface area has high adsorption power of CO<sub>2</sub>, which contributes to reaction acceleration. Specifically, a compound having a specific surface area of 40 to 200 m<sup>2</sup>/g is preferably used, and the use of a compound having a specific surface area of 70 to 150 m<sup>2</sup>/g is further preferable. However, the present invention does not exclude a compound having a specific surface area greater than the above described value.

As a strong acid to be loaded, SO<sub>4</sub><sup>2-</sup> or PO<sub>4</sub><sup>3-</sup> is preferably used. However, any strong acid other than the above can be loaded on the aforementioned carrier. As SO<sub>4</sub><sup>2-</sup> or PO<sub>4</sub><sup>3-</sup>, SO<sub>4</sub><sup>2-</sup> or PO<sub>4</sub><sup>3-</sup> derived from H<sub>2</sub>SO<sub>4</sub>, H<sub>3</sub>PO<sub>4</sub>, (NH<sub>4</sub>)<sub>2</sub>SO<sub>4</sub>, or (NH<sub>4</sub>)<sub>3</sub>PO<sub>4</sub> can be used.

It is preferable that SO<sub>4</sub><sup>2-</sup> or PO<sub>4</sub><sup>3-</sup> of 0.1 to 6% by weight, further preferably 2 to 4% by weight, be loaded on a compound having a solid acid site.

Next, a dimethyl ether modified catalyst in accordance with the present invention is explained by a producing method thereof. A carrier composed of a compound having a solid acid site is prepared as described below.

Specifically, as a carrier, a hydroxide or an oxide

containing at least one kind of zirconium (Zr), aluminum (Al), and titanium (Ti) is prepared. The hydroxide or oxide is obtained by adding a salt of said metal(s) to alkali such as aqueous ammonia and precipitating. At this  
5 time, the mixing ratio of the salt of said metal to alkali is preferably 1:1 to 1:15, further preferably 1:2 to 1:6, at a molar ratio.

In order to render such a precipitate as an oxide, thermal decomposition of the hydroxide or any other  
10 ordinarily used method is carried out. The temperature necessary for the thermal decomposition is 300°C or higher, and is preferably in the range of 400 to 800°C so that the carrier can be held stably at the reaction conditions.

As a carrier, any thus prepared hydroxide or oxide  
15 having a solid acid site can be used. In particular, oxides such as  $\text{ZrO}_2$ ,  $\text{Al}_2\text{O}_3$ , and  $\text{TiO}_2$  are preferably used because catalyst molding can be performed relatively easily and the molding strength can be kept.

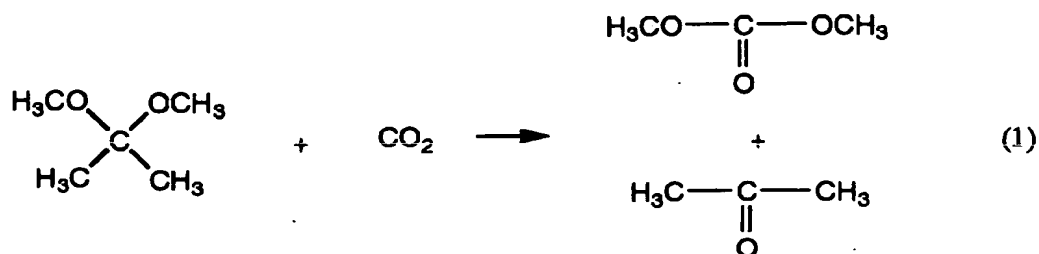
As a precursor of  $\text{SO}_4^{2-}$  or  $\text{PO}_4^{3-}$  that is added to the  
20 carrier, any substance that is turned into  $\text{SO}_4^{2-}$  or  $\text{PO}_4^{3-}$  by calcination at about 500°C may be used. For example, sulfuric acid, ammonium sulfate, phosphoric acid, ammonium phosphate, and the like can be used. The precursor is not limited to above-described substances.

25 As a method for adding a catalyst component, there

can be adopted, for example, a method in which a carrier composed of a hydroxide or an oxide is immersed in aqueous solution containing a precursor of  $\text{SO}_4^{2-}$  or  $\text{PO}_4^{3-}$  at a mol concentration of about 0.01 to 10, preferably about 0.1 to 5, or is impregnated with aqueous solution to perform treatment. Next, the catalyst obtained by loading a strong acid is calcined. As the calcination condition, it is preferable that after being treated by the precursor of  $\text{SO}_4^{2-}$  or  $\text{PO}_4^{3-}$ , the catalyst be stabilized by the calcination at a temperature of about 400 to 800°C, preferably about 450 to 700°C, for about 0.5 to 30 hours. The catalyst prepared by the above-described method is a metal oxide containing  $\text{SO}_4^{2-}$  or  $\text{PO}_4^{3-}$  and having an acid site.

Next, a synthesis system for dimethyl carbonate using the catalyst for dimethyl carbonate synthesis in accordance with the present invention is explained. Hereunder, a synthesis scheme for dimethyl carbonate is shown.





At the system start time, equal mols of acetone dimethyl acetal and CO<sub>2</sub> in a supercritical state are required as starting materials. As shown in Formula (1), these materials are allowed to react in the presence of the catalyst for dimethyl carbonate synthesis in accordance with the present invention. By this reaction, intended dimethyl carbonate is yielded at a conversion rate of about 20 to 40%, and at the same time, acetone (dimethyl ketone), which is a side reactant, is yielded.

The critical temperature of carbon dioxide is 31.3°C, and the critical pressure thereof is 7.3 MPa. Therefore, in the reaction using the catalyst in accordance with the present invention, CO<sub>2</sub> can be allowed to react with acetone dimethyl acetal under the condition that CO<sub>2</sub> is in a supercritical state. Specifically, it is preferable that the reaction using the catalyst of the present invention be carried out in the range of reaction

temperature of 150 to 200°C and reaction pressure of 10 to 30 MPa.

Thus, when the system is brought into normal operation, acetone dimethyl acetal can be produced from acetone, which is a by-product. Therefore, acetone dimethyl acetal itself need not be supplied separately. By supplying CO<sub>2</sub> in a supercritical state, which is used for the reaction expressed by Formula (1), and methanol, which is used for the reaction expressed by Formula (2), to the system as occasion calls, the reactions expressed by Formulas (1) and (2) are allowed to proceed, by which dimethyl carbonate can be obtained.

The above-described system using the catalyst for dimethyl carbonate synthesis in accordance with the present invention is a clean system capable of obtaining dimethyl carbonate by using CO<sub>2</sub> and methanol, which are low in cost and easy to supply, as raw materials without discharging a by-product that requires separate treatment, so that this system can be used preferably for the industrial mass production of dimethyl carbonate.

More specifically, the synthesis system for dimethyl carbonate using the catalyst for dimethyl carbonate synthesis in accordance with the present invention can be incorporated as a part of a methanol synthesis plant. In this case, CO<sub>2</sub> can be obtained from the off-gas of product

methanol and methanol synthesis source gas.

### Examples

Next, the present invention is explained in more  
5 detail with reference to examples. The following examples  
do not restrict the present invention.

#### <Example 1> Preparation of Catalyst 1

Two kilograms of zirconium oxychloride ( $\text{ZrOCl}_2 \cdot 8\text{H}_2\text{O}$ )  
was dissolved in 15L of pure water, and aqueous ammonia  
10 was dripped gradually until the pH value became 10 while  
the solution was agitated. After the yielded zirconium  
hydroxide had been aged for 24 hours, it was filtered,  
cleaned, and dried in vacuum at  $110^\circ\text{C}$ , by which about 800  
g of white powder carrier of zirconium hydroxide was  
15 obtained. After this carrier of composite metal hydroxide  
had been immersed in sulfuric acid of 0.5 mol  
concentration, and excess sulfuric acid had been filtered,  
the carrier was dried and calcined at  $600^\circ\text{C}$  for three  
hours, by which Catalyst 1 ( $\text{SO}_4^{2-}/\text{ZrO}_2$ ) was obtained.

#### 20 <Example 2> Preparation of Catalyst 2

Two kilograms of zirconium oxychloride ( $\text{ZrOCl}_2 \cdot 8\text{H}_2\text{O}$ )  
was dissolved in 15L of pure water, and aqueous ammonia  
was dripped gradually until the pH value became 10 while  
the solution was agitated. After the yielded zirconium  
25 hydroxide had been aged for 24 hours, it was filtered,

cleaned, and dried in vacuum at 110°C, and was further calcined at 650°C, by which about 650 g of white powder carrier of zirconium oxide was obtained. After this carrier of composite metal hydroxide had been immersed in phosphoric acid of 1 mol concentration, and excess phosphoric acid had been filtered, the carrier was dried and calcined at 550°C for three hours, by which Catalyst 2 ( $\text{PO}_4^{3-}/\text{ZrO}_2$ ) was obtained.

<Example 3> Preparation of Catalyst 3 and Catalyst 4

Catalysts 3 and 4 were obtained by the same method as that in Example 2, except that titanium tetrachloride and aluminum nitrate were used respectively, in place of the zirconium oxychloride used as a carrier component in Example 2.

<Example 4> Preparation of Catalyst 5

Two kilograms of commercially available zirconium oxychloride ( $\text{ZrOCl}_2 \cdot 8\text{H}_2\text{O}$ ) and 2.33 kg of aluminum nitrate ( $\text{Al}(\text{NO}_3)_3 \cdot 9\text{H}_2\text{O}$ ) were dissolved in 15L of pure water, and aqueous ammonia was dripped gradually until the pH value became 10 while the solution was agitated. After the yielded zirconium hydroxide/aluminum hydroxide ( $\text{Zr}(\text{OH})_x \cdot \text{Al}(\text{OH})_x$ ) composite metal hydroxide had been aged for 24 hours, it was filtered, cleaned, and dried in vacuum at 110°C, by which about 1500 g of white powder carrier was obtained. After this carrier of composite metal hydroxide

had been introduced in phosphoric acid of 2 mol concentration, and excess phosphoric acid had been filtered, the carrier was dried and calcined at 550°C for three hours, by which Catalyst 5 ( $\text{PO}_4^{3-}/\text{ZrO}_2 \cdot \text{Al}_2\text{O}_3$ ) was  
5 obtained.

<Example 5> Preparation of Catalyst 6

Catalyst 6 was obtained by the same method as that in Example 4, except that titanium tetrachloride was used, in place of the zirconium oxychloride used as a carrier  
10 component in Example 4.

<Comparative example> Preparation of Comparative catalyst 1 and Comparative catalyst 2

The carrier of zirconium oxide prepared in Example 2 was used as Comparative catalyst 1, and the carrier of  
15 aluminum oxide prepared in Example 3 was used as Comparative catalyst 2.

The compositions and surface areas of these catalysts are given in Table 1.

20 Table 1 Composition and surface area of catalyst

Catalyst number	Catalyst composition	Specific surface area ( $\text{m}^2/\text{g}$ )
Catalyst 1	$\text{SO}_4^{2-}/\text{ZrO}_2$	45
Catalyst 2	$\text{PO}_4^{3-}/\text{ZrO}_2$	55

Catalyst 3	$\text{PO}_4^{2-}/\text{TiO}_2$	61
Catalyst 4	$\text{PO}_4^{3-}/\text{Al}_2\text{O}_3$	102
Catalyst 5	$\text{PO}_4^{2-}/\text{ZrO}_2 \cdot \text{Al}_2\text{O}_3$	88
Catalyst 6	$\text{PO}_4^{3-}/\text{TiO}_2 \cdot \text{Al}_2\text{O}_3$	95
Comparative catalyst 1	$\text{ZrO}_2$	52
Comparative catalyst 2	$\text{Al}_2\text{O}_3$	105

<Example 6> Evaluation of dimethyl carbonate synthesis activity of catalyst

The dimethyl carbonate synthesis activity was evaluated by using catalysts prepared in Examples 1 to 5 and Comparative example. The synthesis reaction was carried out using an autoclave under the conditions of 185°C, 30 MPa, and 0.5 g of catalyst amount by mixing carbon dioxide ( $\text{CO}_2$ ), methanol, and acetone dimethyl acetal (ADA), which are raw materials, so that the amount is 21.3 g, 3.2 g, 3.2 g, respectively. The result was evaluated by the concentration of dimethyl carbonate (DMC) yielded by the reaction. The yield ratio of dimethyl carbonate in the case of reaction using Catalysts 1 to 6 and Comparative catalysts 1 and 2 are given in Table 2 to compare the catalyst activity.

Table 2 Reaction conditions for evaluating catalyst activity and comparison of activity

	Catalys 1	Catalys 2	Catalys 3	Catalys 4	Catalys 5	Catalys 6	Comparative catalyst 1	Comparative catalyst 2
DMN yield ratio <sup>1)</sup>	0.30	1.00	0.92	0.80	0.95	0.86	<0.01	<0.01
Temperature	185 °C							
Pressure	30 MPa							
Catalyst amount	0.5 g							
Raw materials	21.3 g/ 3.2 g/3.2 g							

1) Expressed by relative value, taking the yield for the catalyst having the highest activity as 1.

5 From the results of Table 2, it was found that the yields of the catalysts in accordance with the present invention increase significantly as compared with the yield of the conventional catalysts.

The catalyst for dimethyl carbonate synthesis in  
10 accordance with the present invention can be used under the supercritical condition of CO<sub>2</sub>, and dimethyl carbonate can be obtained from acetone dimethyl acetal and CO<sub>2</sub> in a supercritical state at a high conversion rate of about 27%, which is five times or more of the conventional conversion  
15 rate. Also, because a solid acid catalyst is used, the

catalyst for dimethyl carbonate synthesis in accordance with the present invention has no toxicity unlike the conventional case where phosgene or CO is used, can be handled easily, and further is advantageous in terms of  
5 cost.

Furthermore, by using such a catalyst for dimethyl carbonate synthesis, a system can be constructed in which methyl carbonate is yielded by way of acetone dimethyl acetal as a reaction intermediate product and by using  
10 methanol and CO<sub>2</sub> as raw materials. Because such a system has been incapable of being realized by the conventional method for the reason of reaction equilibrium, the clean system of the present invention is very advantageous in producing dimethyl carbonate industrially. Further, this  
15 system can be incorporated in a system for obtaining a raw material for a fuel cell, and therefore is also promising for performing highly efficient operation of a fuel cell.